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Table 17.3 Solubilities of Ionic Compounds* aq = aqueous (dissolves in water); s = solid (does not dissolve in water)

Ions	Acetate	Bromide	Carbonate	Chlorate	Chloride	Fluoride	Hydrogen Carbonate	Hydroxide	Iodide	Nitrate	Nitrite	Phosphate	Sulfate	Sulfide	Sulfite
Aluminum	s	aq	aq	aq	aq	s		s	—	aq	aq	s	aq	—	—
Ammonium	aq	aq	aq	aq	aq	aq	aq	—	aq	aq	aq	aq	aq	aq	aq
Barium	aq	aq	s	aq	aq	s	—	aq	aq	aq	aq	s	—	—	s
Calcium	aq	aq	s	aq	aq	s	—	s	aq	aq	aq	s	—	—	s
Cobalt(II)	aq	aq	s	aq	aq	—	—	s	aq	aq	s	aq	s	s	s
Copper(II)	aq	aq	s	aq	aq	aq	—	s	aq	aq	s	aq	s	s	s
Iron(II)	aq	aq	s	aq	aq	s	—	s	aq	aq	s	aq	s	s	s
Iron(III)	—	aq	—	aq	aq	s	—	s	aq	aq	s	aq	—	—	—
Lead(II)	aq	s	s	aq	aq	s	—	s	aq	aq	s	s	s	s	s
Lithium	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	s	aq	aq	aq
Magnesium	aq	aq	s	aq	aq	s	—	s	aq	aq	aq	s	aq	—	aq
Nickel	aq	aq	s	aq	aq	aq	—	s	aq	aq	s	aq	s	s	s
Potassium	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq
Silver	s	s	s	aq	s	aq	—	—	s	aq	s	s	s	s	s
Sodium	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq	aq
Zinc	aq	aq	s	aq	aq	aq	—	s	aq	aq	s	aq	s	s	s

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