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Experiment 25 Report Sheet
Calorimetry

Date: 11/17/18 Lab Sec: _____ Name: Kim Sorensen Desk No: _____

③ Specific Heat of a Metal

	Trial 1	Trial 2
1. Mass of metal (g)	63.852	63.852
2. Temperature of metal (boiling water) (°C)	93.5°C	95°C
3. Mass of calorimeter (g)	4.032	4.032
4. Mass of calorimeter + water (g)	23.238	23.238
5. Mass of water (g)	19.206	19.206
6. Temperature of water in calorimeter (°C)	19°C	25°C
7. Minimum temperature of metal and water from graph (°C)	28.5°C	24.5°C

8. Instructor's approval of graph: _____

Calculations for Specific Heat of a Metal

	Trial 1	Trial 2
1. Temperature change of water, ΔT (°C)	74.5	70.0
2. Heat gained by water (J)	6180	580
3. Temperature change of metal, ΔT (°C)	74.5	70.0
4. Specific heat of metal, equation 25.5 (J/g·°C)		
5. Average specific heat of metal (J/g·°C)		

*Show calculations for Trial 1 using the correct number of significant figures.

1) ΔT (°C) = $93.5 - 19 = 74.5$

2) $q = 19.206 \text{ g} (4.18 \text{ J/g} \cdot \text{°C}) (74.5)$

3) $\Delta T = 74.5^\circ$

4) $q = \frac{(m \Delta T)_{\text{water}}}{(m \Delta T)_{\text{metal}}}$

$q = 6180$

Experiment 25 301

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