

# Download File PDF 2013 Ap Government Multiple Choice Answers

#Jenny



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#Markus Jensen



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#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

- Name \_\_\_\_\_  
AP Chemistry Summer Assignment 2011 - 2012 School Year  
Part 4 Assignment continued  
Multiple Choice Questions  
You MUST show all work for any problems involving math!
- Which one of the following elements has been selected as the current atomic weight standard?  
a. O  
b. C  
c. H  
d. Na
  - Bromine is composed of two isotopes. One of the isotopes, Br-79, makes up 49.7% of the total while the other, Br-81, makes up 50.3% of the total. Calculate the atomic mass of Br.  
a. 80.0  
b. 80.5  
c. 81.7  
d. 79.9
  - Four beakers containing potassium nitrate dissolved in water are allowed to evaporate to dryness. Beakers 1 through 4 contain 2.3, 1.91, 5.985, and 0.52 g of dry potassium nitrate. How many moles of potassium nitrate were recovered after the water evaporated?  
a. 0.104 moles  
b. 0.212 moles  
c. 0.500 moles  
d. 2.35 moles
  - How many atoms of uranium (U) are present in 1 nanogram of uranium?  
a.  $3.5 \times 10^{21}$   
b.  $5.0 \times 10^{21}$   
c.  $2.5 \times 10^{21}$   
d.  $5.0 \times 10^{21}$
  - Calculate the percent composition of hydrogen in sucrose ( $C_{12}H_{22}O_{11}$ ).  
a. 6%  
b. 10%  
c. 33%  
d. 39%
  - Two of the three forms of vitamin B6 are pyridoxine ( $C_8H_{13}NO_3$ ) and pyridoxamine ( $C_8H_{13}N_2O_3$ ). Calculate the respective percentage of nitrogen in each compound.  
a. 8.0%, 17%  
b. 9.3%, 19%  
c. 9.2%, 17%  
d. 18%, 9.2%
  - Determine the molecular formula of a compound that contains 26.7% P, 12.1% N, 61.2% O, and a molecular weight of 580 g/mol.  
a.  $(PNO)_2$   
b.  $(PNO)_3$   
c.  $(PNO)_4$   
d.  $(PNO)_5$

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